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(54) Title: SALINE SOLUBLE INORGANIC FIBRES

(57) Abstract

Disclosed is use of a vitreous inorganic fibre in the knowledge that it has a composition meeting the criterion that the calculated sum of the free energies of hydration of the compounds that would or could be present at equilibrium (on the basis of knowledge, informed belief or reasonable assumption) is more negative than -10 kcal/100 grams of compositions are saline soluble.

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SALINE SOLUBLE INORGANIC FIBRES

This invention relates to saline soluble, non-metallic, amorphous, inorganic oxide, refractory fibrous materials.

Inorganic fibrous materials are well known and widely used for many purposes (e.g. as thermal or acoustic insulation in bulk, mat, or blanket form, as vacuum-formed shapes, as vacuum formed boards and papers, and as ropes, yarns or textiles; as a reinforcing fibre for building materials; as a constituent of brake blocks for vehicles). In most of these applications the properties for which inorganic fibrous materials are used require resistance to heat, and often resistance to aggressive chemical environments.

Inorganic fibrous mate: als can be either glassy or crystalline. Asbestos is an norganic fibrous material one form of which has been strongly implicated in respiratory disease.

It is still not clear what the causative mechanism is that relates some asbestos with disease but some researchers believe that the mechanism is mechanical and size related. Asbestos of a critical size can pierce cells in the body and so, through long and repeated cell injury, have a bad effect on health.

Whether this mechanism is true or not regulatory agencies have indicated a desire to categorise any inorganic fibre product that has a respiratory fraction as 'hazardous', regardless of whether there is any evidence to support such categorisation. Unfortunately for many of the applications for which inorganic fibres are used, there are no realistic substitutes.

Accordingly there is a demand for inorganic fibres that will pose as little risk as possible (if any) and for which there are objective grounds to believe them safe.

A line of study has proposed that if inorganic fibres were made that were sufficiently soluble in physiological fluids that their residence time in the human body was short; then damage would not occur or at least be minimised. As the risk of asbestos-linked disease appears to depend very much on the length of exposure this idea appears reasonable. Asbestos is extremely insoluble.

As intercellular fluid is saline in nature the importance of fibre solubility in saline solution has long been recognised. If fibres are soluble in physiological saline solution then, provided the dissolved components are not toxic, the fibres should be safer than fibres which are not so soluble. The shorter the time a fibre is resident in the body the less damage it can do. H. Förster in 'The behaviour of mineral fibres in physiological solutions' (Proceedings of 1982 WHO IARC Conference, Copenhagen, Volume 2, pages 27-55(1988)) discussed the behaviour of commercially produced mineral fibres in physiological saline solutions. Fibres of widely varying solubility were discussed.

International Patent Application No. W087/05007 disclosed that fibres comprising magnesia, silica, calcia and less than 10 wt% alumina are soluble in saline solution. The solubilities of the fibres disclosed were in terms of parts per million of silicon (extracted from the silica-containing material of the fibre) present in a saline solution after 5 hours of exposure. The highest value revealed in the examples had a silicon level of 67 ppm. In contrast, and adjusted to the same regime of measurement, the highest level disclosed in the Förster paper was equivalent to approximately 1 ppm. Conversely if the highest value revealed in the International Patent Application was converted to the same measurement

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regime as the Förster paper it would have an extraction rate of 901,500 mg Si/kg fibre - i.e. some 69 times higher than any of the fibres Förster tested, and the fibres that had the highest extraction rate in the Förster test were glass fibres which had high alkali contents and so would have a low melting point. This is convincingly better performance even taking into account factors such as differences in test solutions and duration of experiment.

International Patent Application No. WO89/12032 disclosed additional fibres soluble in saline solution and discusses some of the constituents that may be present in such fibres.

European Patent Application No. 0399320 disclosed glass fibres having a high physiological solubility.

Further patent specifications disclosing selection of fibres for their saline solubility are European 0412878 and 0459897, French 2662687 and 2662688, PCT WO86/04807 and WO90/02713.

The refractoriness of the fibres disclosed in these various prior art documents varies considerable. The maximum service temperature of any of these disclosed fibres (when used as refractory insulation) is up to 815°C (1500°F).

Service temperature for refractory insulation is definable in many ways but to be consistent with the above mentioned International Patent Applications this application shall mean by service temperature that temperature at which the fibre shows acceptable shrinkage (maximum of 5% linear shrinkage after exposure to temperature for 24 hours) and at which the fibre has not appreciably suffered through excessive sintering or softening.

There is a demand for physiologically soluble fibres having a service temperature of greater than 815°C,

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particularly for such fibres having a service temperature above 900°C.

Testing for physiological solubility and safety can be done by inhalation studies on, e.g. rats. However such studies are extremely time consuming and costly. A study can take of the order of $2\frac{1}{2}$ years from start and can easily cost f1 million per study. A cheaper alternative is to test for solubility in physiological or like fluids in vitro.

Testing of an inorganic fibre for solubility in physiological solutions is not so time consuming, but there is currently no way of predicting which systems will produce such soluble fibres. Therefore anyone seeking to find such soluble fibres has to work on a trial and error basis assisted by what is commonly known as 'chemical intuition' but is equally commonly known as 'a hunch'. Such trial and error testing is laborious and time consuming. Further, once a fibre is found that is soluble there is no guarantee that it will be usable at useful service temperatures.

Accordingly there is a demand for a method of predicting whether a fibre will have a reasonable solubility in physiological solutions, and further there is a demand that such a test should preferably give an indication as to expected service temperature.

Shrinkage of inorganic refractory fibres occurs through two mechanisms; the first is viscous flow of the fibre material. Most inorganic refractory fibres are glasses and so may be defined as liquids having an exceedingly high viscosity (but still liable to flow). By their nature fibres are elongate and so have a high surface area per unit volume. As the reduction of surface area is a means of reducing the surface energy of a material, when the glass becomes fluid enough it will flow so as to reduce surface area. This flow results in a coarsening and shortening of the fibres and so to

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shrinkage, and in the extreme results in disruption of the fibres into separate particles.

The second mechanism leading to shrinkage is that at elevated temperatures glasses may crystallise to form one or more crystal phases. Usually these crystal phases have a smaller molar volume than the glasses from which they crystallise and so shrinkage results. Some fibres are known for which the molar volume of the crystalline form exceeds that of the glass (for example Al₂O₃/SiO₂ glassy fibres may crystallise to form mullite crystals). In these cases the expansion due to crystallisation may oppose the shrinkage caused by viscous flow.

If shrinkage through viscous flow occurs at a much lower temperature than crystallisation then the crystallisation may not be able to compensate for such shrinkage.

There is a demand for a fibre in which both viscous flow and crystallisation occur at as high and as similar a temperature as possible, and preferably in which the expansion due to crystallisation closely matches the shrinkage due to viscous flow so that the net effect is as close to zero shrinkage as possible.

When used as refractory insulation inorganic refractory fibres are used in several forms. The fibres may be supplied as a bulk material, but in this form the fibres are difficult to handle for many applications. Alternatively the fibre may be supplied as a blanket. Blanket fibre is generally made by a process of sucking fibre from air onto a conveyor to form a blanket. Because the fibres tend to be aligned parallel to the conveyor surface they can separate easily. Accordingly the blanket fibres are secured together by adding a binder to lock the fibres together, or by needling the blanket, or both. In needling needles are passed through the thickness of the blanket to push and draw fibres to lie transverse to the

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blanket and so tie the fibres together. Because binders are usually resins, such as phenolic resins, they burn off on first firing. There is a desire to reduce the amount of such binders used both because of possible health implications in handling, and because the combustion products may affect the strength of the fibres. Thus needled blanket is usually preferred.

The fibres may also be supplied as blocks, generally made from assembled layers of inorganic fibre blanket.

For some fibres needling is not possible. Crystalline fibres—are generally too brittle to stand the stresses involved. For the fibres known in the industry as glass fibres (which are generally used for low temperature applications) the amount of 'shot' (unfiberised glass particles) present is generally too high to allow needling as the shot damages the needles. There is no needled blanket on the market that has a maximum service temperature in the range 900°C-1200°C. There are needled blankets having a higher maximum service temperature but these use expensive fibres in comparison with other fibres usable (with the aid of binders) as blanket in the temperature range 900°C-1200°C.

Accordingly there is a demand for needled fibre blanket formed from inexpensive materials, being soluble in saline solutions, and having a maximum service temperature in the range 900°C-1200°C.

As stated previously refractory oxide fibres are made by several methods all of which involve the formation of a melt of oxides and the subsequent fiberisation of the melt by e.g. spinning or blowing.

The melt of oxide material is often formed by electrical discharge melting of the constituent raw materials. The applicants, in manufacture of a CaO/MgO/SiO2 refractory oxide

fibre encountered problems due to the necessity of handling CaO. These problems were discovered to be due to the moisture content of CaO as commercially available. One of the problems of use of CaO is the outgasing that results upon melting and this led at the least to a porous melt pool which caused fluctuations in the melt current; in the extreme the outgasing was explosive. Additionally use of CaO appeared to cause accelerated attack on the melt electrodes. Also CaO is a difficult and corrosive material to handle.

Accordingly there is a need for a process that minimises the use of CaO.

Accordingly the present invention provides the following features both independently and in combination:-

- A. Use of a vitreous inorganic fibre in the knowledge that it has a composition meeting the criteria either: that the calculated sum of the free energies of hydration of the compounds that would or could be present at equilibrium (on the basis of knowledge, informed belief or reasonable assumption) is more negative than -10 kcal/100grams (-418.4kJ/kg) of composition; or that the percentage of non-bridging oxygens is more than 30%. Such compositions tend to be saline soluble.
- \underline{B} . Use of such a vitreous inorganic fibre in the knowledge that it has a composition meeting the criterion that the ratio of glass modifiers to network formers is less than a critical value (for SiO_2 based compositions the critical value is 1). Such compositions tend to be glass formers.
- <u>c</u>. The invention also encompasses fibres selected by adopting such criteria as a test for solubility and glass formation.
- D. Use as saline soluble fibres having a shrinkage of less

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than 3.5% when exposed to 1000°C for 24 hours and/or 800°C for 24 hours, of vitreous fibres having a composition comprising (in weight %):-

$$SiO_2$$
 >58% - (for MgO =< 10%) and SiO_2 >58% + 0.5(%MgO -10) - (for MgO >= 10%) (SiO₂ preferably being below 70%) CaO 0% - 42% MgO 0% - 31.33% Al₂O₃ 0% - <3.97%

and being essentially free of fluxing components such as alkali metals and boron oxide.

 \underline{E} . In one such usage the first crystalline material resulting on crystallisation has the crystal structure of diopside and has the composition consisting essentially of:-

Component	Composition A
	Weight percent
sio ₂	59-64
Al ₂ 0 ₃	0-3.5
CaO	19-23
MgO	14-17

F. In a second such usage the first crystalline material resulting on crystallisation has the crystal structure of wollastonite/pseudowollastonite and has the composition consisting essentially of:-

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Component	Composition B		
	Weight percent		
SiO ₂	60-67		
Al ₂ O ₃	0-3.5		
CaO	26 - 35		
MgO	4-6		

- G. The fibres used in such manner may further be used as needled blankets.
- $\underline{\text{H}}$. Preferably the fibres of the general composition and compositions A and B mentioned above have a SiO_2 content (expressed as a weight percentage of the constituents SiO_2 , CaO and MgO) of greater than 60%.
- I. The present invention further provides a method for the manufacture of refractory oxide fibres containing calcium and silicon by the formation of an oxide melt containing calcium and silicon characterised in that all or part of the calcium and all or part of silicon is provided by a calcium silicate.

The invention is illustrated by way of example in the following description and with reference to the drawings in which:-

- Fig.1 is a three-axis phase diagram indicating the crystalline phases in the system SiO₂/CaO/MgO (*Phase Diagrams for Ceramists*, The American Ceramic Society, 1964) a key to this diagram is at the end of the specification;
- Fig.2 is a three-axis composition plot of the projection onto the ${\rm SiO}_2/{\rm CaO/MgO}$ phase field of compositions comprising ${\rm SiO}_2$, CaO, MgO and ${\rm Al}_2{\rm O}_3$;
- Fig.3 is a temperatur /time plot of the firing regime used in a series of cyclic exposure tests of experimental

compositions;

Fig.4 is a plot of log (total solubility) v calculated free energy of hydration for a series of fibres.

Fig.5 is a plot of log (total solubility) v % non-bonding oxygens for a series of fibres (see below).

A series of fibres were made of the compositions shown in Table 1. These fibres were melt spun by using a vertical spinning system of the type known for making inorganic fibres. Also shown in Table 1 are the compositions of some comparative commercially available inorganic oxide fibres and glass fibres.

TABLE 1

	$^{20}3$	\mathtt{SiO}_2	CaO	Mg0	Zro ₂	
SW-A	3.3	59.3	20.5	15.5		
SW-Al	1.1	63.7	20.5	15.2		
SW-A2	0.8	60.8	21.4	15.4	-	
SW-B1	2.3	. 65.3	26.8	5.7	-	
SW-B2	1.3	66.9	27.5	5.2	_	
SW-B3	1.0	60.0	34.0	4.4	-	
			COMPAR	ATIVE	EXAMPLES	
	Al_2O_3	sio ₂	CaO	MgO	zro ₂	MST
CRBT	46.5	53	0.04	0.01	-	1260°C
CWBT	40.6	49.5	5.50	4.00		870 °C
CHBT	49.7	35.1	0.04	0.01	14.7	1425°C
Glass	15.2/	53.7/	21.1/	1.3/		+ 5.9-6.2% B ₂ O ₃
Fibre	15.5	57.5	21.8	1.6		0.11-0.12% TiO ₂
1 1220						0.46% Na ₂ O
						0.32-0.33% K ₂ O
Needled	3.7	60.5/	8.1	4.0		+ 2.85-2.95% B ₂ O ₃
Glass	J.,	60.0	7.9		_	13.5% Na ₂ O
Fibre						1.0% K ₂ 0
						

[MST = Maximum Service Temperature (oxidising atmosphere)]

The fibres SW-A, SW-A1, SW-A2, SW-B1, SW-B2 and SW-B3 were tested for solubility by the following method.

The fibre was first chopped in the following manner. 2.5 g of fibre (deshotted by hand) was liquidised with 250 cm³ of distilled water in a domestic Moulinex (Trade Mark) food blender for 20 seconds. The suspension was then transferred to a 500 cm³ plastic beaker and allowed to settle after which as much liquid as possible was decanted and the remaining liquid removed by drying in an oven at 110°C.

The solubility test apparatus comprised a shaking incubator water bath, and the test solution had the following composition:-

Compound	Name	<u>Grams</u>
NaCl	Sodium cnloride	6.780
NH ₄ Cl	Ammonium chloride	0.540
NaHCO ₃	Sodium bicarbonate	2.270
Na ₂ HPO ₄ .H ₂ O	Disodium hydrogen	0.170
2 4 2	phosphate	
Na ₃ C ₆ H ₅ O _{7P} .2H ₂ O	Sodium citrate	0.060
3 0 3 7 2	dihydrate	
H2NCH2CO2H	Glycine	0.450
H_2SO_A s.g. 1.84	Sulphuric acid	0.050

The above materials were diluted to 1 litre with distilled water to form a physiological-like saline solution.

0.500 grams \pm 0.0003 grams of chopped fibre was weighed into a plastic centrifuge tube and 25 cm³ of the above saline solution added. The fibre and saline solution was shaken well and inserted into the shaking incubator water bath maintained at body temperature (37°C \pm 1°C). The shaker speed was set at 20 cycles/minute.

After the desired period (usually 5 hours or 24 hours)

the centrifuge tube was removed and centrifuged at ≈ 4500 revs/minute for approximately 5 minutes. Supernatant liquid was then drawn off using a syringe and hypodermic needle. The needle was then removed from the syringe, air expelled from the syringe, and the liquid passed through a filter (0.45 micron cellulose nitrate membrane filter paper [WCN type from Whatman Labsales Limited]) into a clean plastic bottle. The liquid was then analysed by atomic absorption using a Thermo Jarrell Ash Smith - Hiefje II machine.

The operating conditions were:-

ELEMENT	WAVELENGTH (nm)	BAND WIDTH	CURRENT (MA)	FLAME (Nitrous Oxide + Acetylene)
Al	309.3	1.0	8	Fuel Rich
SiO ₂	251.6	0.3	12	11 11
CaO	422.7	1.0	7	Fuel Lean
MgO	285.2	1.0	3	11 11

The procedure and standards adopted for determining the above elements were as set out below.

 SiO_2 can be determined without dilution up to 250 ppm concentration (1 ppm = lmg/Litre). Above this concentration an appropriate dilution was made volumetrically. A 0.1% KCl solution (0.1g in 100 cm³) was added to the final dilution to prevent ionic interference. NB If glass apparatus is used, prompt analysis is necessary.

From a stock solution of 1000 ppm pure ignited silica (99.999%) (fused with Na_2CO_3 at 1200°C for 20 minutes in a platinum crucible (0.2500g $SiO_2/2g$ Na_2CO_3) and dissolved in dilute hydrochloric acid (4 molar) made up to $250cm^3$ with distilled water in a plastic volumetric flask) the following

standards were produced:-

STANDARD (PPM SiO ₂)	STOCK SOLUTION (cm3
10.0	1.0
20.0	2.0
30.0	3.0
50.0	5.0
100.0	10.0
250.0	25.0

Add 0.1% KC1 to each standard before making to 100cm3.

Aluminium may be measured directly from the sample without dilution. Standards of 1.0, 5.0 and 10.0 ppm Al may be used. For calibration readings are multiplied by 1.8895 to convert from Al to ${\rm Al}_2{\rm O}_3$.

A standard Al atomic absorption solution (e.g. BDH 1000 ppm Al) was bought and diluted using an accurate pipette to the desired concentration. 0.1% KCl was added to prevent ionic interference.

Calcium may require dilutions on the sample before determination can be carried out (i.e. \times 10 and \times 20 dilutions). Dilutions must contain 0.1% KC1.

A standard Ca atomic absorption solution (e.g. BDH 1000 ppm Ca) was diluted with distilled water and an accurate pipette to give standards of 0.5, 4.0 and 10.0 ppm. 0.1% KC1 is added to prevent ionic interference. To convert readings obtained from Ca to CaO a factor of 1.4 was used.

Magnesium may require dilutions on the sample before determinations can be made (i.e. \times 10 and \times 20). Add 0.1% KC1 to each dilution. To convert Mg to MgO multiply by 1.658.

A standard Mg atomic absorption solution (e.g. BDH 1000 ppm Mg) was diluted with distilled water and an accurate

pipette to give standards of 0.5, 1.0 and 10.0 ppm Mg. 0.1% KCl was added to prevent ionic interference.

All stock solutions were stored in plastic bottles.

The results of the tests are indicated in Table 2.

TABLE 2

BODY FLUIDS SOLUBILITY

(ppm)

	SiO ₂		Ca	1 0	M	MgO		
	5h	24h	5h	24h	5h	24h		
SW-A	98	120	63	56	33	66		
SW-Al	83	141	32	70	21	70		
SW-A2	130	202	43	73	100	177		
SW-B1	58	77 -	10	38	5	9		
SW-B2	64	121	27	- 55	5	10		
SW-B3	138	192	80	46	8	21		

Fibres with the best solubility (SW-A2 and SW-B3) were then tested, after annealing at varying temperatures, and compared with the comparative examples of Table 1. The results are shown in Table 3.

It can be seen that for the SW-A2 fibre, with increasing annealing temperature, the silica solubility drops progressively. In contrast the SW-B3 composition shows no loss in solubility up to 800°C and although a reduction in solubility is shown above that temperature it is not as dramatic as for SW-A2. Despite this difference in solubility it is to be noted that only the needled GF fibre shows a

comparable silica solubility and that material melts at 700°C.

TABLE 3

3	37	37	3	89	130
3	56	27	43	60	108
5	53	17	30	43	87
7	3	3	2	11	21
5	69	7	22	22	100
1 :	150	12	22	55	130
1	90	3	7	24	144
8	40	3	3	17	60
0	8	6	3	5	3 -
				•	
6	10	7	3	4	0.3
4	17	5	3	5	7
7	34	8	15	66	85
1	26	7	10	19	37
6	16	7	6	8	9
	33 3 5 7 5 1 1 8 0 6 4 7 1 1	3 37 3 56 5 53 7 3 6 69 1 150 1 90 8 40 0 8 6 10 4 17 7 34 1 26	24hrs 5hrs 3 37 37 3 56 27 5 53 17 7 3 3 5 69 7 1 150 12 1 90 3 8 40 3 0 8 6 6 10 7 4 17 5 7 34 8 1 26 7	See 24hrs 5hrs 24hrs 3	3 56 27 43 60 5 53 17 30 43 7 3 3 2 11 5 69 7 22 22 1 150 12 22 55 1 90 3 7 24 8 40 3 3 17 0 8 6 3 5 6 10 7 3 4 4 17 5 3 5 7 34 8 15 66 1 26 7 10 19

[[]The Glass Fibre and Needled Glass Fibre had the compositions shown in Table 1.]

The user is primarily concerned with the solubility of the fibre as received as it is in this condition that most handling occurs; as received both SW-A2 and SW-B3 fibres have extremely high solubility. Even after exposure to 800°C and 1000°C these fibres have solubilities much higher than other high temperature use fibres.

To investigate the reasons underlying the difference in solubilities after high temperature annealing between the SW-A2 and SW-B3 fibres qualitative X-ray diffraction was done on the fibres. The results are indicated in Table 4 and it can be seen that the SW-B3 fibre forms pseudowollastonite and wollastonite, whereas the SW-A2 fibre forms diopside. It appears therefore that the crystalline diopside has a lower solubility in physiological saline solution than the crystalline pseudowollastonite and wollastonite material precipitated from the SW-B3 fibre.

TABLE 4

<u>Sample</u>	Condition	<u>Oualitative XRD</u>
SW-A2 SW-A2	600°C, 48 hours 800°C, 48 hours	Amorphous Amorphous with small amount Diopside
SW-A2	1000°C, 48 hours	Diopside
SW-B3	600°C, 48 hours	Amorphous
SW-B3	800°C, 48 hours	Amorphous
SW-B3	1000°C, 48 hours	Pseudowollastonite & Wollastonite

Various of the fibres were then tested for their shrinkage characteristics. Table 5 shows the results of Shrinkage tests on all the test fibres and on some of the comparative fibres. These results were obtained by proposed ISO standard ISO/TC33/SC2/N220 (equivalent to British Standard

BS 1920, part 6,1986) with some modifications to account for small sample size. The method in summary comprises the manufacture of vacuum cast preforms, using 75g of fibre in 500cm^3 of 0.2% starch solution, into a 120 x 65mm tool. Platinum pins (approx 0.1-0.3mm diameter) were placed 100 x 45mm apart in the 4 corners. The longest lengths (L1 & L2) and the diagonals (L3 & L4) were measured to an accuracy of $\pm 0.01 \text{mm}$ using a travelling microscope attached to a steel rule with a vernier scale. The samples were placed in a furnace at temperature and left for 24 hours. The shrinkage values are given as an average of the 4 measurements.

TABLE 5

LINEAR SHRINKAGE (%)

(24h at temperature)

					·	
Temp.°C	SW-A	SW-A1	SW-A2	SW-Bl	SW-B2	SW-B3
					,	. 1
730	1.45		1.43	1.02	0.22	
870	,			0.41		
900			1.07	•		1.07
1000		1.04	1.3	0.51	0.6	1.1
1100	3	0.71	1.8		0.73	2.2
Maximum Service Temperatu °C	850 re	1050	1050	1050	1050	1000

It can be seen that in SW-A, SW-A1, SW-A2, SW-B1, SW-B2 and SW-B3, owing to the increase in molar volume on crystallisation, the linear shrinkage at the maximum service temperature is less than 3.5%.

Table 6 shows the results of a further series of shrinkage tests made in the same way.

TABLE 6

Sample	Measurement Direction c.f. Roll Direction	Test Temperatures C	Linear Shrinkage Range	es % Mean
SW-A2	Parallel	850	1.1-1.4	1.2
SW-A2	Perpendicular	850	0.7-1.5	1.3
SW-A2	Parallel	900	0.5-1.1	0.9
SW-A2	Perpendicular	900	1.9-4.5	3.0
SW-A2	Parallel	1000	0.5-2.9	1.3
SW-A2	Perpendicular	1000	1.7-2.9	2.2
SW-A2	Parallel	1100	0.7-1.5	1.0
SW-A2	Perpendicular	1100	1.0-2.6	1.8
SW-B3	Parallel	900	1.6-1.8	1.7
	Perpendicular	900	1.4-2.4	2.1
SW-B3	Parallel	1000	1.6-2.3	1.9
SW-B3	Perpendicular	1000	1.0-2.3	1.7
SW-B3	Parallel and	1100	Complete	Meltir
	Perpendicular		(Lantern remnant)	type

To ascertain the applicability of these tests to long term usage a series of cyclic shrinkage tests were undertaken on the materials and the heating schedule used for these cyclic

tests is shown in Figure 3.

The results of the tests are shown in Tables 7 & 8 (the two figures given for SW-B3 are due to slight differences in chemical analysis [the fibre at the end of a production run of fibre tends to have slightly differing composition to that at the beginning of a production run of fibre]).

As a further comparison with the above discussed materials a melt was made comprising 55% SiO₂, 29.9% CaO and 18.6% MgO. Fibres made using this composition had a maximum service temperature of 700°C and melted at 800°C.

As these results were encouraging the applicants conducted a further and extensive series of tests, concentrating on the SW-A2 and SW-B3 compositions, to ascertain the reproducibility of these results and the boundaries of the useful compositions.

Table 9 (three pages) below gives the compositions of a series of melts, ranked on silica content, and showing the shrinkage figure after exposure to 1000°C for 24 hours (1st column) and 800°C for 24 hours (2nd column). shrinkages were measured by the same method as the shrinkages given above but measurements were made with a travelling microscope with a digital linear scale accurate to \pm 5 μ m. can clearly be seen that all fibres with a silica content of less than 58% have a shrinkage at 1000°C of greater than 3.5% save two (B3-3 and 708). These fibres, together with some fibres with a silica content of greater than 58% although showing a reasonable figure at 1000°C, show a very poor figure at 800°C. Compositions with an SiO2 content of greater than 70% appear to fiberise poorly. This may be because such compositions have two liquids in the melt as may be appreciated from Fig.1

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TABLE 7
CYCLIC SHRINKAGE (LINEAR)

(왕)

Product	1000°C	1100°C	24h at
710000			1000°C
No. cycles	58	42	
CRBT	2.0	2.7	1.9
CWBT	15.0	13.3	12.1
SW-A2	0.33	2.0	1.3
SW-B3	1.00	1.67	1.1
SW-B3	0.33	0.67	1.1

Accuracy: + or - 0.33%

TABLE 8
CYCLIC SHRINKAGE

(%)

		LINEAR SHRINKAGE	. *	THICK SHRIN	
PRODUCT	1000°C	1100°C	24 h at 1000°C	1000°C	1100°C
No. cycles	104	100		104	100
CRBT CWBT SW-A2 SW-B3 SW-B3	1.47 14.4 1.5 1.73 1.47	3.1 15.2 2.1 1.63 1.77	1.9 12.1 1.3 1.1	0.47 38.63 8.58 7.24 7.02	11.19 32.14 8.75 7.57 7.16

Accuracy: + or - 0.3%(%)

IE	
00	

			Anal	Analysed Co (Weight	Compositight %)	ıtıons	·		Shrinkag at °C	, C	Rationali Compositi		sed
بر	Sio2%	Ca0%	Mg0%	A12038	Na208	K20%	Fe2038	Zro2%	1000	800	Sio2%	Ca0%	Mg0%
80	3.0	2.0	17.1	-	0.0	0.0	0.1	7	F		:	2	7.
ø	3.4	12.4	10.0	-	0	0.	0.1	2	0.7	*	•	_	_
2	3.0	6.3	19.6	N	0	0.0	0.1	E,	<u>-</u>		÷		6
2	3.3	23.4	9.0	3	E.	0.09	0.5	7					0
S	2.2	12.6	12.3	7	0	0.	0.1	2		*	•	щ е	2
22	71.48		16.34	0.33	0.10	<0.05	0.20	0.83	1.1		73.6	9.6	16.8
6	1.2	4.4	22.3	4	0	0.	0.1	7		*1.5	å	•	2
27	1.1	6.7	19.6	N	٥.	0	0.1	0	•	;	٠.		ö
8	8.0	18.7	7.0	4	~	0	0.1	۲.	•	*	÷	•	
7	4.0	11.5	14.5	7	0	0	0.1	ល	•	*		2	•
33	0.0	4.6	22.8	~	۲.	۰.	0.1	•	F		÷.	4.	щ
8	4.6	23.2	3.7	4	ų.	0	0.1	4.			'n	•	m.
	9.2	15.1	12.7	0	~	0	0.1	ᅻ.			Ξ.	ů.	•
9	8:7	24.9	1.7	9	.	4	0.2	۰.		•	ċ	ė	<u>.</u>
9	8.6	8.1	19.2	7	0	۰.	0.2	o.	•	•		ω.	•
27	8.5	20.9	7.0	7	۳.	•	0.1		•	•	-	•	7
	8.3	17.4	12.6	0	₽.	•	0.3	٥.	•	•	<u>.</u>	7	'n
20	8.1	11.5	16.5	4		0.05	0.2	σ.	1.1	1.0		8	•
31	7.6	8.4	21.7	e.	0	۰.	0.1	œ.	•	•	ດ	œ.	8
31	7.5	27.7	0.4	4.	7.	~	0.2	•	•		ċ	ص	•
61	7.5	24.9	3.6	4.	L.	٥.	0.1	œ.	•		ċ	•	•
17	7.2	26.6	1.8	7	4	٦.	0.2	ເນ	•		o.		
13	9.9	14.8	16.0	7	•	•	0.1	ຫຼ	•		α.	ທີ	•
22	6.1	21.2	9.3	r.		0	0.1	ru.	•	0.8	φ.	'n	9
10	6.1	16.2	15.8	4.	0.	•	0.0	•	•	٠	Ŀ.	ė.	•
15	5.8	29.8	1.7	4.	4	•	0.1	ι.	•			ċ	•
	5.7	25.6	8.1	0		•	0.2	0	•		Ġ.	Ď.	•
מ	5.6	18.7	13.7	Γ.	٦.	٠.	0.0	Ξ.	•	1.0	ė.	o.	•
4 A	5.5	25.8	4.8	0	.,	Ξ.	_ 		•	•		1	
	ıc	7 2	. 71	r	-	٦	_	7			67.7	7.2	14.8

TABLE 9 (continued)

	-		Analys (W	0 0	a Composit lght %)	itions		·	Shrinka at °(kage °c	Rationa Composi	11; [t]	sed
Melt	sio2%	CaO%	Mg0%	A12038	Na20%	K20%	Fe2038	Zr028	1000	800	$\mathrm{Sio}_2 ^{\$}$	Ca0%	Mg0%
	2	27.1	6.9	0	~	0.0>	0.2	4			2	7	7.0
	1	24.9	5	5	4	0.0	0.1	9			68.1	26.1	
-	0	27.2	53	0	7	<0.0>	0.0	0		•	9	7	ري س
1	8	9.9	26.2	N	0	<0.0>	0.1	8		3.9	9	9	٠
	-	13.7	19.9	, CJ	_	<0.0>	0.1	۲		•	ហ	•	0
1	ן ה	31.9	0.3	9	4	0.0	0.1	N			ė	ب	0
1		8	24.8	N	~	<0.0>	0.1	4	_	7.1	ហ	о О	•
12	0	23.2	9.3	Ŋ	m	0.0	0.1	G.	_	,	ė.	٠	•
ונו	7.	25.4	4.6	0	N	0.1	0.1	ល	-	4.3		,	
E	3.6	16.0	18.2	4	_	0.0	0.1	0.0		•	ທີ	ġ.,	
6	9.	21.4	12.9	4	n	0.1	0.1	0	-	•	64.9	21.9	13.2
֚֚֚֡֝֟֝֝֝֝֟֝ <u>֚</u>	ເລ	16.5	18.0	e.	_	<0.0>	0.0	0	į.	ς.	4	ė.	œ
ļ	2	24.8	4.5		N	0.1	0.1	۲.		•			
	9	29.7	5.4	0	_	0.0>	0.1	0		•	4	•	ົ. ເດ
	9	20.9	15.2	7	0	0.0>	0.1	9		•		÷	•
1 .		10.9	24.1	۵,	Ξ.	0.0	0.1	0	- 4			;	4
ا .		30.6	2.0	щ	Ψ.	0.0	0.1	ο,	•	0.9	ъ.	2	'n
1 1	, 0		11.98	0.44	0.25	0	0.10	0.23	1.4		63.7	24.0	12.3
ı.		28.1	7.5	7		×0.0	0.1	٠.	•		ю	ω.	٠
֚֚֚֚֚֚֚֚֚֝֞֞֝֟֝֟֝֟֝֟֝֟֝֟֝֟֝֓֓֓֓֓֓֓֓֟֝֟֝֓֓֓֟֝֓֓֓֓֟֝֓֓֓֓֓֓֓֓		33.2	2.3	Ψ.		0.	0:	41	•		3	4	2
10	9	25.5	9.7	E.1	.,	0.		·	•	Н	3	•	٠
	1.3	13.6	22.7		Γ.	×0.	0.1	٠.	•	•	N	m,	m ı
2	1.3	31.0	5	٠.		×0.	-	٠	•		'n	-	٠ ا
		19.7	14.5		٦.	0° 0°	0	Ψ.	•	3.0	4	0	•
3-2	60.8	32.3	0.4	Γ.			-		•	•		ı	
2-1	60.7	25.3	11.6		٠.	<u>.</u>	0.1	٦.	•		62.2	25.9	היד
3-1	60.3	32.2	3.0	٠.		<u>.</u>	; -	٠,	-	•	,	•	(
2	60.3	24.2	13.2	٠.	•	0 V	<u>.</u>	٠,	•	1:1	61.7	24.8	13.5
3-9	60.	34.4	2.5	•	٠.	<u>.</u>	<u></u>	7	, i	ή,	· N	٠ د	N (
A2-23	-	18.	18.7	٦.	•	0		٠.	•	•	Ξ.	<i>2</i> 0	•
1													

TABLE 9 (continued)

ysed Compositions Shrinkage Rationalised (Weight %) at °C Compositions	Al ₂ O ₃ \$ Na ₂ O\$ K ₂ O\$ Fe ₂ O ₃ \$ ZrO ₂ \$ 1000 800 SiO ₂ \$ CaO\$ MgO\$.79 0.54 0.12 0.23 0.82 0.5 62.3 33.8 3.	.50 0.06 <0.05 0.18 0.11 12.	.05 0.08 0.17 2.1 60.0 34.5 5.	.91 0.33 0.15 0.21 0.41 43.	.48 0.39 0.08 0.14 <0.05 2.3 61.4 29.1 9.	.92 0.43 0.14 0.23 0.17 0.9 61.8 37.5 0.	.05 0.27 <0.05 0.40 <0.05 0.7 1.2 59.5 32.2 8.	.21 0.18 <0.05 0.08 0.24 0.3 1.4 61.3 32.3 6.	.54 0.20 0.07 0.19 <0.05 4.7 60.6 19.0 20.	.13 0.23 <0.05 0.16 <0.05 1.9 1.6 60.1 35.9 4.	.27 0.20 <0.05 0.10 0.17 1.6 1.9 59.9 33.9 6.	.12 <0.05 <0.05 0.23 <0.05 34.2 59.7 4.0 36.	.15 0.24 <0.05 0.19 <0.05 24.7 59.4 35.6 5.	.91 0.57 0.13 0.22 <0.05 30.9 59.8 33.9 6.	.06 0.23 <0.05 0.12 <0.05 4.4 3.4 59.0 36.3 4.	.73 0.51 0.11 0.16 0.08 30.3 59.5 38.0 2.	.62 0.44 0.11 < 0.05 < 0.05 39.8 26.6 58.8 31.5 9.	.66 0.51 0.10 0.19 <0.05 36.2 59.2 27.3	.74 0.18 <0.05 0.14 0.05 41.7 57.9 24.4 17.	.00 0.06 <0.05 0.13 0.05 39.3 58.2 24.3 17.	.50 0.08 <0.05 0.14 0.05 34.9 57.7 25.1 17.	.58 0.45 0.09 0.38 1.46 0.3 34.7 58.0 37.4 4.	.05 0.31 <0.05 0.12 <0.05 1.1 8.7 55.9 43.3 0.	
	Zroz	ω.		۲.	4.	0.	ı.	0.	7	0:	0.		0.	0.	0	•	0.	•	0	•	•	•	4	۰.	C
	e ₂ 0 ₃	.2		0	7	ı.	7	4.	0.	4	۲.	۲.	7	۲.	2	٦.		•		۲.	٦.	۲.	ω.	۲.	•
ions	20	۲.	0	0	۲.	0	٦.	0	0	0	0.0	0.0	0.0	0.0	۲.	•		۲.		•	0	•	•	0	100
mposit: %)	a ₂ 0	.5	0		۳.	ω.	4.	2	۲.	7	7	2	0	2	ស	~	נו	4.	ស		•	0	4	<u>س</u>	
ysed (Weig	1203	7.	ເນ	ļ.	0	4.	9	0.	~	٠	۲.	7	۲.	r.	6	0	.7	9.	9	.7	0	ນ	5	0	•
Anal	Mg0&	8	9	5.44	9	7	9	2	۲.	.7	3.9	0	ď	4.8	0	9	4.	4	3.0	7.3	0	6.8	4.4	.7	•
	CaO%	2.5	1.6	34.34	1.6	8.1	6.0	2.0	0.9	8.4	5.0	3.1	3.9	4.6	2.7	5.3	6.6	0.5	6.2	3.9	3.6	4.4	6.0	2.7	,
	Sio2%	0.1	9.8	59.82	9.5	9.5	9.4	9.0	8.8	8.7	8.6	8.3	7.7	7.7	7.5	7.5	7.3	6.9	6.9	8.9	6.5	6.1	9	5.2	
	Melt	B3-1	32	692	3-3	B3-25	3-2	14	969	A2-19	86	694	765	099	B3-20	12	3-2	3-2	A2-18	7.1	734	973	B3-3	08	. 1

*Poor fibres containing a lot of shot. All other constituents < 0.1%. Too poor to test for solubility or shrinkage

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sed	Mg0%	17.6	.	9		5	16.8	N	· •		15.0	m ı	m (13.1	<u>.</u>	$\frac{20.1}{1}$	_	N I	17.2	N	0.0		٠,	16.4	9	9	8.4	80	J.4.C	<	7.4.
	9/0		6	4	m	-	9	2	5	4	0		-	9	. 2	4.	.7	.7	0	9	0	9	<u>,</u>	7	ŗ.	0	9	œ (-		•
nal	CaO	2	12.	ဖ်	24.	13.	o .	4 1		13	2	4	4	12	ဖ	∞ .	21	^	N	ш	ווט	25	. •		·		_	<u> </u>	٠.		1
Ration	1028	80.3		•	•	•	•	•	•			•	-:	•				_	_	~	$\overline{}$	0	\circ	on. I	^	œ	~	9	Ø	ŧ	/./0
	0	8	5 2	_	_	8 7	5	5	7 7	17	2 7		7	5	<u>-</u>	=	0	5	6	~	4		<u>. </u>	0	2		_	4	0	16	44
ities	Mg		īĊ			4	~	10	9	<u>س</u>	7								9			~								_	
111t	CaO		45				36						9	4	7	ന	 5	n	<u>е</u>	~	រប	44	4	4	<u>~</u>	Ψ	6.1	u,	<u>~</u>		
fiqnlos mdd	sioz	-				10		თ	10	10	თ	-	14	12	13	18	12	11	16	2	11	163	16	15	12	17	E	7	님	ω,	1
	2%	~	2	C	7	~	æ	C	0	-	Œ J	O	4	_	_	01		_	٠.	۳.	٦.	0.65	-:	•	-				•	•	• 1
	Zro	4	6	o.	•	4	ö	<u>-</u> i	i	•	۲,	ri	0	<u>.</u> -i	ri	<u> </u>	0	0	0	0	<u></u>	<u> </u>	0	0	0	0	0	<u>~</u>	0	0	1
	3%	1	Н	٦	2	0	0	n	<u>ر</u>	9	[3	7	7	11	24	21	17	33	21	15	22	.15	23	10	60	18	18	23	10	16	14
	Fe ₂ 0	0	ö	Ö	0	0	o	Ö	0	0	Ö	0	o	Ö	0	ò	ò	0	0	0	0	0	0	0	0	0	0	0	0	0	<u> </u>
In .	90	10	0	0	0	0	0	0	0	0	0	U	U	J	_	_	_	_	_	_	٠.	.07	٠.	٦.	٠.	Ξ.	-	-		•	•
lon	K20	0	٥٠.	0	0	<0×	. 0×	, 0 0	Ö	0	00	o	Ö	00	o	0	0	0	0	0	0	ò	0	0	0 >	0	0	0	0	0	<u> </u>
11t	%	95	25	25	37	0.5	10	05	08	23	0.5	11	37	25	38	0.5	35	11	19	60	40	37	45	05	90	33	45	24	16	28	13
d Compositions ight %)	Na 2	0	×0.	0	0	\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \	0	· 0×	0	0	, 0 2	0	0	0	0	\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \	-	6	0	0	0	0	0	0	0	0	0	_	0	0	<u>.</u>
col	3%																					45									22
Analysed (Wel	A120	6	0	c			0	0	c	c		ċ	c	Ċ	Ċ	_	· C	\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \	, c	- C	Ç	0	_	-	· c	_	- C	\ 0 \	0	7	0
nal	96	-	0	· ·	ט כ) C	, C	C	v	, C) a					• •	, 4				9			•	•	•	• •			
A	Mgo	17	10.	9		2.0	9	22	6	, ,	7	2,0	, ~		-	4 0	1 6	, ,	1 -	2 5	1 0	· "	_	1 4	1 -	1 0	, -	4 CC	7	4	14
	9/0		7	• ~	7	ד ע) L	7	. L	, [. K	ט כ	ם כ	, L.	4 0		7 0			., `		6	. ~	. ~	•	: :	: -		•	. ~	•
	Cao	6		!	כי	. C L	3 6	4		ָ מ	֓֝֝֜֜֜֜֜֜֜֝֓֜֜֜֜֜֜֓֓֓֓֓֜֜֜֜֜֜֓֓֓֓֜֜֜֜֓֓֡֓֜֜֡֓֡֡֡֜֜֡֓֡֡֡֡֡֡	• • •	֓֜֜֝֓֓֓֓֓֓֓֓֓֓֓֓֓֓֓֓֡֓֓֓֓֡֓֓֡֓֡֓֡֓֡֓֓֡֓֡֓֡֓֡	ה		7 0	9	7 6	;;	-i 0	0 7	3 6	2 2	7 7	1 -	7 6	4 0	4 C	3 5	2 1 2	16
	9%																					ט מ									
	sio	1	•	•	•	•	•	•	•	: _	: -	: -	: -	: -	: -	: .	ά.		8	0 0	0 / 0	7 7	, נ ט	0	0 1	٥	0 1	ט ט	פע	ט ע ט	9 6
	بد	c	۹ ۲	4 0	2 0	י ר	2 5	4 5	9 (9 0	٧,	٦,	? .	! ነ	٠,	ተ የ		ν,	ͺ (? (ט נ	7 6	- 	⊣ •	⊣ « !	7 (7 0	ת ת	1 1	P
	Me1																														A2

TABLE 10 (continued)

	9													
			Anal	ysec (We)	(ght %)	tions			Solubili ppm	bilit ppm	ies	Ration Compos	ali iti	sed
Melt	Sio2%	ca0%	. %0£M	A12038	Na208	K20%	Fe ₂ 0 ₃ \$	${ m Zr0}_2 \$$	Sio2	CaO	MgO	Si02%	ca0%	Mg0%
718	5.2	27.1	6.9	0.0>	2	0.0>	0.2	4	01	47		ູນ.	7	•
B3-14	65.11	24.91	5.54	0.58	0.43	0.09	0.19	0.61	158	67	27	68.1	26.1	ۍ 8
21	0.0	27.2	5.3	0.0	4	<0.0>	0.0	0	15	68		9	7.	•
2-3	4.8	6.6	26.2	0.2	0	<0.0>	0.1	æ	c	11		9	ė.	•
A2-21	1.1	13,7	19.9	0.3	4	<0.0>	0:1	۲.	16	28		ິນ	•	ö
3-3	4.1	31.9	0.3	9.0	4	0.0	0.1	2	16	83		9	3	0
2-3	4.1	8.8	24.8	0.2	7	0.0>	0.1	4.	9	18	106	ນ	٠ 0	•
3-2	4.0	23.2	9.3	0.5	ω.	0.0	0.1	ų.	20	44	31	9	•	•
3-5	3.7	25.4	4.6	3.9	N	0.1	0.1	r.	4	43	11			
2-3	3.6	16.0	18.2	0.4	Ξ.	0.0	0.1	0	10	33	52	ъ.	9	ж •
2-9	3.6	21.4	12.9	1.4	ω.	0.1	0.1	0	19	67	99	64.9	21.9	13.2
2-1	3.5	16.5	18.0	0.3	٦.	<0.0>	0.0	٥.	16	47	99	4.	9	œ
1	3.2	24.8	4.5	5.7	~	0.1	0.1	۲.	~	7	26			
23	2.6	29.7	5.4	0.0>	ㄷ.	<0.0	0.1	•	14	62	17	4.		ທ
57	2.6	20.9	15.2	0.2	0	<0.0>	0.1	•	18	62	,	3.	;	•
2-2	2.3	10.9	24.1	0.3	₹.	0.0	0.1	۰.	16	35	105	3	-	4
3.	2.3	30.6	2.0	0.5	ø.	0.0	0.1	ດຸ	12	41	7	2	8	ς.
2-7	1.9	23.3	11.9	0.4		<0.0	0.1	7	17	59	63	٠. ص	4	٠
25	1.8	28.1	7.5	0.1	ų.	0.0 0.0	0.1	•	18	76	35	63.4	28.9	7.7
3-1	1.7	33.2	2.3	0.6	r.	0.1	0.2	ເຄ	10	28	15	3	4	2
B3-24	1.6	25.5	9.7	0.5	۳.	0.1	0.1	٠.	78	37	57	٠ .	· ·	•
2-2	1.3	13.6	22.7	0.3	⁻:	0.0	0.1	•	=	27	87	3	m	ب ا
22	1.3	31.0	5.2	0.0	~	0.0	0.7	0	8	81	23	'n	_	<u>ب</u>
924	1.3	19.7	14.5	2.5	0	0.0>	0.0	9		36	37	4.	0	٠
3-2	9.0	32.3	0.4	4.1	۳.	0.1	0.1	Ξ.	u,	19	Ŋ			
2-1	0.7	25.3	11.6	0.2	Ξ.	0.0	0.1	Ξ.	72	57	43	62.2	25.9	11.9
3-1	0.3	32.5	9.6	1.7	<u>.</u>	0.1	-	0	Ξ	72	16			
2	0.3	24.2	13.2	0.2	Τ.	0.0 0.0	0.0		7	54	51	61.7	24.8	13.5
3-9	0.2	34.4	2.5	0.7	ιΩ	0.7	<u>.</u>	7	1,	ω	ഗ	2	ល	8
A2-23	2	18.5	18.7	0.4	٦.	0.0	<u>.</u>	٠.	5	35	47	1.	6	•
1					1	-								

TABLE 10 (continued)

			-								Ì			
		v	Anal	ysed Comp (Weight ?	soc ()	itions			Solubili ppm	bilit ppm	ies	Ration	ali iti	sed
Melt	sio2%	Ca0%	MgO&	A12038	Na 20%	K20%	Fe2038	Zroz&	sio2	cao	Mgo	Sio2%	Ca0%	Mg0%
		200	8	1	ហ	0.1	0.2	æ	1.4	84	17	2.	3	EL,
	. α	22.5	75.6	٠.IU	0	0.0	0.1	Н	Q	41	43	9.19	22.3	16.1
202	. «	34.3	5 4	י ר	_		0.08	0.17	140	67	23		4.	•
֓֞֞֞֜֜֞֜֞֜֜֓֓֓֓֓֓֓֓֓֓֓֓֓֓֓֓֓֓֓֡֓֓֡֓֓֓֓֡֜֝֡֓֓֡֓֜֝֡֓֡֓֜֝֡֓֡֓֡֝		31.6	9.0	ש	S.	0.1	0.2	4	m	25	10			
	. נכו	28.1	9.2	4	m	0.0	0.1	0	13	48	333	٠	თ I	
֡֜֝֝֞֜֜֜֜֝֜֜֜֝֝֓֜֜֜֜֝֓֜֜֜֜֜֜֜֜֜֜֜֜֜֜֝֜֜֜֜֝֜֜֜֜֝֡֓֜֜֜֜֝֡֜֜֜֝֡	4	36.0	0.6	Q	4	0.1	0.2	4	20	29	S	•	7.	-
, 4		32.0	8.2	0	N	0.0>	0.4	0	14	73	30		8	_
		30.9	6.1	~	٦.	<0.0>	0.0	ď	91	48	37		'n	٠ و
•		18.4	19.7	ເດ	N	0.0	0.1	0	12	42	99	<u>.</u>	9	-
ָ ע		35.0	3.9	-	Ŋ	0.0>	0.1	0	18	54	19	·	•	_
ח ה	מ	33.1	9	N	C.	<0.0>	0.1	Γ.	13	99	CA.	59.9	33.9	9 9
۱ IC	7	3.0	35.0	٦.	0	<0.0>	0.2	•	2	16	169	·	•	
วน		34.6	4.8	٦.	C	<0.0>	0.1	٠.		69	23	<u>.</u>	•	
"		32.7	9	0,	г.	0.1	0.2	9	18	29	40	· ·	m	
, כ י		35.3	4.6	0	~	0.0	0.7	0.0	16	24	22	<u>.</u>		•
יוני הייני		36.6	2.4		נט	0.1	0.1	٠.	17	25	16	·	m	•
76		30.5	6.6	Ψ.	4	0.1	0.0>	٠.	14	64	41	m.	.i 1	ຸດ ເ
ן ו ו	י טע	26.2	13.0		71	0.1	0.0	٠.	ĭ	47	57	<u>.</u>		•
1 5	י טע	23.0	17.3		7	<0.0	0.7	٠	14	ລີ	73		4.	
• (טע	2 2 2	17.0			<0.0	<u>.</u>	٠.	C.	52	99	œ.	4.	٠.
7 5	֓֞֜֜֜֜֜֓֓֓֓֓֜֜֜֓֓֓֓֜֜֜֜֓֓֓֓֜֜֜֜֓֡֓֡֓֜֜֜֜֡֓֡֓֡֡֡֡֡֓֜֜֜֡֡֡֡֡֡	2 4 4	10			<0.0	<u>.</u>	٦.	7	42	47		ຜ	۲.
? ` (יו טופ	200			,	0.0	0	`.	7	14	18	œ		•
7 (ה ה	2 5				0>	0	٦.	7	31	N	'n	e,	Ö
27	54.68	24.04	19.66	0.20	0.17	×0.	0	٠.	7	55	62	ល	4	•
+ .							-			-				

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There are several anomalies, namely compositions B3-6A, A2-25, A2-24, A2-23, B3-2A, B3-3A, A2-19, and 932. All of these have an SiO_2 content of > 58% but a high shrinkage.

On the assumption that the minimum silica level for satisfactory shrinkage varies with MgO content the applicants have determined that fibres with a silica content (in weight percent) that fail to meet the following expression do not have satisfactory shrinkages at either or both 800°C and 1000°C:-

$$SiO_2$$
 >58% - (for MgO =< 10%) and SiO_2 >58% + 0.5(%MgO -10) - (for MgO >= 10%)

The applicants have further found that the ${\rm Al}_2{\rm O}_3$ content is important. From their studies it appears that the maximum ${\rm Al}_2{\rm O}_3$ content lies somewhere between 2.57% and 3.97%. The applicants have found that with increasing alumina levels the first material to crystallise is calcium aluminate and this possibly forms a liquid phase that assists flow and hence shrinkage.

Table 10 shows, for the same compositions as Table 9, 24 hour solubilities for each major constituent. It can be seen that all of the compositions have high solubilities.

As mentioned above use of CaO in forming calcium containing fibres is inconvenient and can be hazardous. The applicants investigated use of mixed oxide materials that would avoid the handling of CaO. A fibre was made by admixture of magnesia with silica and wollastonite (CaSiO₃).

The raw materials used to make the melt comprised:-

Pennine Darlington Heavy Magnesia (#200)

MgO	-	92.60%
CaO	-	1.75%
Fe ₂ O ₃	- ,	0.04%
sio ₂	_	0.20%
Cl	-	0.25%
so ₃	-	0.70%
LOI	-	4.50%

Partek's Finnish Low Iron Wollastonite (#200)

(U.K. agent - Cornelius Chemical Co., Romford, Essex)

sio ₂	-	51.80%
CaO	-	44.50%
MgO	-	0.80%
A1 ₂ 0 ₃	-	0.60%
Fe ₂ O ₃	_	0.30%
Na ₂ O	-	0.10%
к ₂ 0	-	0.05%
TiO2	-	0.05%
s	-	0.02%
MnO	_	0.01%
P	_	0.01%
F	· -	0.01%
LOI	-	1.70%

Hepworth Mineral's Redhill T washed silica sand

sio₂ - 99.0% min.

These constituents were mixed as 78.65% Wollastonite; $19.25\% \, \mathrm{SiO}_2$; and 3.6% MgO. This gave 0.4 - 0.5% of the final melt as $\mathrm{Al}_2\mathrm{O}_3$.

It was surprisingly found that in producing a melt using these constituents the current requirements were only two-thirds that for the use of the raw oxides.

Fibre was produced by blowing (although spinning and other methods can be used). 2 runs were performed with different blowing conditions.

Chemical analysis was undertaken by the Analytical Department at the applicant's subsidiary Morgan Materials Technology (hereinafter called M²T) using wet chemical techniques. Fibre diameters were measured using M²T's Galai particle analyser, with shape analysis software. Typically 40,000 fibres were analysed for each run.

The first result of note was the speed of melt reaction when using wollastonite as compared with lime. Also the current was seen to be very stable throughout the growth of the melt. If the current was lost whilst pulling the electrodes apart the current could be restored simply by pushing them back together again. This was not possible with the runs using lime.

Chemical Analysis

	<u>Others</u>	<u>CaO</u>	MgO	<u>Al₂O</u> 3	<u>sio</u> 2	ZrO ₂	<u>Na₂0</u>	<u>Total</u>
Runl	0.7	32.6	3.8	0.8	60.1	0.8	0.5	99.3
Run2	0.7	32.5	3.8	0.8	60.1	0.8	0.6	99.3

Runs 1 and 2 indicate the respective x-ray fluorescence analyses for each run.

Shrinkage Results (1000°C for 24 hours)

	<u>L1</u>	<u>L2</u>	<u>L3</u>	L4	AV.	Std.Dev.
Run 1	0.9	0.2	0.4	0.6	0.5	0.3
Run 2(A)	1.0	-0.2	0.7	0.6	0.5	0.5
Run 2(B)	0.5	0.2	0.0	0.4	0.2	0.2

Solubility Results (ppm)

	<u>Ca0</u>	MgO	<u>sio</u> 2
Run 1(5 hr)	67	10	95
Run 1(24hr)	84	17	146
Run 2(5hr)	39	7	72
Run 2(24hr)	73	17	186

Fibre diameters

	Mean	Median	100%<	%>5µm	%<1µm
Run 1	5.1µm	$3.4 \mu \mathrm{m}$	$30 \mu m$	33%	13%
Run 2	4.1µm	$2.7 \mu m$	25μm	25%	19%

Accordingly it appears to be the case that by using what are cheaper ingredients than the pure oxides one can obtain a fibre that has as high a performance as using purer oxides and at much improved energy costs and safety. It is to be noted that this feature of the invention is not limited to saline soluble fibres and any oxide fibre that contains both calcium and silicon can advantageously be made with a calcium silicate, wollastonite being merely an example of such a silicate.

The previous description is directed towards high temperature usage of particular saline soluble fibres. The following is directed towards the prediction and use of saline soluble fibres. A series of fritted glass melts were made of the compositions shown in Tables 11A and 11B and quenched in water. Solubilities of the various components of the quenched melt were measured by the previously described method of atomic absorption. The solubilities were normalised to a specific surface area of 0.25m²/gram of fibre.

The fr e energy of hydration was calculated by normalising the chemical analysis to 100 weight %; making the assumption that simple silicat s (MSiO $_3/M_2$ SiO $_3$) are present

and calculating the free energy of hydration contribution of each species: and summing to get the total free energy of hydration. The data in Tables 11A & 11B is also presented in Fig.4. It can be seen that the fibres lie on a generally straight line bar four groups of materials that will be explained below.

Table 11 shows in each column the following:-

Fibre reference
Composition
Molar ratios
Moles/100 grams of melt
Species assumed (see below)
Calculated free energy of hydration of assumed species
(kcal/100 grams)(see below)
Calculated free energy of hydration of assumed species
(kJ/kg) (see below)
Solubility data (# indicates not measured [see below])
Specific surface area
Normalised solubility data
log normalised solubility

The base data on which calculation of the free energy of hydration was done is set out in Tables 12 which indicates free energies of hydration taken from the literature in the units kcal/mol and kJ/mol.

Table of free Energy of Hydration Values for Stilcate Helts

TABLE 11

																					٢,		_			_			_	_		_	- 1		_	-	_	_	_		i	
Log(norm solub)	1.63				1.82			18	39.				1.84			1	.,			17.7		- 35	0.00			00	3			0.78	_			06.0	_			2.45		انهجم		
2 <u>}</u>		_			3			ļ	•				ŝ			0.5	A 5 7			165			•			a	• —			0								280				
	0.30		_	j	0.39			<u> </u>	95.0				0.40			1	 			0.45			9.20			5	70.0			0.52				0.50				0.41				
e	-	•	~	2	88	- :	3 !	<u>-</u>	ς.	- ·	£	티	Ş	- :	3	- -	· ;	ò	269	Ţ	253	762	•	•	^ <	<u>\</u>	= -		2	-	~	•	2	۵	~	٠.	2	ĸ	2	355	260	
Oxide (pp	i	A1203	Sioz	Total	CaO	A1203	2015	Total	CaO	A1203	2018	Total	CaO	1:02	2105	Total	510	2015	Total	Ca0	Sios	Total	Hn20	A1203	2015	1000	080	At Cub	Total	NgO	A1203	5102	Total	Sro	A1203	\$102	Total	Sro	CoO	Sio2	Total	
Free Energy of Hydration (kJ/kg)	-306-0	31.6	.219.7	-492.1	-322.8	30.6	198.9	.491.1	-307.3	30.6	-223.0	-409.5	5.927-	-45.5	7.011	9.113.	3.709	-21.7	.626.1	.563.5	11.2	.552.3	.367.5	20.7	4.69	4.722	7.152	64.9	7.5%	.169.0	44.5	153.5	0.65	-300.2	9.07	77.9	-181.7	-346.1	-253.4	3,3	-596.2	
free Energy of	-		-5.3	-11.8	-7.7	0.7	8.7-	-11.8	.7.3	7.0	-5.3	-11.9	7.11.	-1.1	2.6	6.0-	7.51-	. 5.0.	-14.9	.13.5	0.3	-13.2	.8.8	1.2	2.1	-5.5	0.9.		0,0	7	-	3.7	1.2	-7.2	1.0		6.4.		-6.		.14.3	
_		7	10.5		4.70	2.35	3.63		4.56	2,37	4.07		7.07	0.63	1.65		5.92	0.30	_	8.36	0.48		3.05	3,90	3.82		3.73	3.43	?	73.5	27.	3,	_	2.94	3, 12	-	<u>}</u>	2	_	_	_	
Species	1010	41201	2	}	Casios	A1203	CaO		Cosio3	A1203	CaO		Casio3	CuD	1102		SrSio3	SiO		Cosio3	\$102	•	He25103	A1203	5005		Costos	A1203	2015	1010	TOCAL	C CO	1	Sesios	A1204	2013	3016	100	74501	\$102	;	
Holes	-	177	107.0		0.479	0.235	0.363	_	0.456	0.237	0.407		0.707	0.083	0.165		0.592	0.030		0.836	0.048		0.305	0.390	0.382		0.373	0.345	0.358		0.620	0.546	;	702 U		_		9		_	_	
Holes	1	20.0	157 0	-	0.842	0.235	0.479		0.863	0.237	0.456		0.790	0.165	0.707		0.622	0.592		0.836	0.884		0.305	0.390	0.687		0.373	0.345			0.230	0.346		0 206		0.5	0.020	15	7,5	9.270	;	
	Hationalized		27.6		47.2	24.0	28.8	100.0	48.4	24.2	27.4	100.0	44.3	13.2	42.5	0.001	64.4	35.6	100.0	6.95	53.1	100.0	18.9	39.8	41.3	0.00t	50.9	35.2	43.9	0.001	10.3	7.		2.00.		2.5	27.7	2		5.5	6.5.0	200
8	ī	7.75		88	5.6.5	23.7	28.4	98.6	47.8	23.9	27.1	98.8	42.7	12.7	41.0	7.96	63.1	34.9	98.0	1 97	52.3	7 80	9	19.7	41.1	9.66	21.0	35.4	44.2	9.00	20.0	33.7	23.0	2	6,4.0	31.1	36.8	1.74	35.0	21.0	43.0	4,.0
E .	Oxide	00 2	20214	John	2	A1203	2015	lotol	003	A1203	205	Total	000	1102	\$102	fotal	5	5102	Total	5	Sin2.	1010	4n20	A1 203	5102	Total	CaO	A1203	2015	Total	W30	A1203	2015		o s	A1203	205	0	Sro	9 ;	2105	10181
Helt Code	1	CAS10(B) - A			CAS10(8) - 8				CAS10/8) - C				CISCA				Srs103			Cacina			MASSAN				CAS	ŗ			HAS(B)				SAS(A)				SCS(A)			

lable of free frergy of Aydration Values for Silicate Helts

TABLE 11 Continued

Helt Code	اد	٤ſ		Holes	Moles	Species	Notes	free Energy of	free Energy of	anjos	Solubility	8.S.A	Hormal Ized	Log(norm	
<u> </u>	o K	٦	Rationalized		in 100g		in ikg	Hydration (kcal/1009)	Hydration (kJ/kg)	Onide	(ucdc)	m2/9	Solubility	solub)	٠.
SMS(A)	S ₂	55.0	9.95	0.546	0.296	srsios	2.96	-7.2	-302.2	SrO	25	97.0	38	1.58	
	8	1.2	1.2	0.030	0.030	Myd	0.30	-0.5	-8.2	OG#	~			!	
	A1203	23.7	7.72	0.239	0.239	A1203	2.39	0.7	31.1	A1 203	~				
	S i 02	17.3	17.8	0.296	0.250	Sro	2.50	÷.3	.181.0	5102	· ^				
	Total	97.2	100.0					-11.0	.460.3	Total	\$			-	
SHAS(B)	Sro	55.0	\$5.9	0.539	0.401	Srsio3	4.01	9.6.	7.607-	5	2	0.39	56	1.98	
	₩g0	4.8	6.4	0.122	0.122	NyO	1.22	9.0-	-33.2	MgO	5				
	A1203	8.×	12.1	0.148	0.148	A1203	1.48	0.5	19.2	A1203	-				_
	2018	23.7	24.1	0.401	0.138	Sro	1.38	-2.4	6.66-	\$ 102	25				
	Total	98.3	100.0					-12.5	-523.3	Total	871				
SHS(A)	SrO	33.0	32.9	0.318	0.318	Srsios	3.18	-7.8	-324.7	Sro	147	0.39	212	2.33	
_	MgO	12.5	12.5	0.310	0.310	MgS103	3.10	-4.3	.180.4	MgO	59		!		
	A1203	1.0	0	0.010	0.010	A1203	0.10	0.0	1.3	A1203	0				
	\$102	53.8	53.6	0.892	997.0	5102	2.64	1.5	61.8	5005	123				
	Total	100.3	100.0					-10.6	.445.0	Total	331				
SAS(C)	Sro	26.0	56.2	0.542	0.401	Srsio3	4.01	8.6-	7.607	Sro	2	0.38	25	1.72	
	A1203	19.7	19.7	0.193	0.193	A1203	1.93	9.0	3.1	ALZOS	-	•			
	S 102	24.0	24.1	0.401	0.141	SrO	1.41	7.2.	-102.1	sin2	•				_
	Total	2.66	100.0					-11.6	-486.4	Yotal	٤	*			
CHAS(A)	003	35.0	34.6	0.617	0.293	CaSio3	2.93	1.4.	-197.5	9	=	0.30	72	1.38	
	N ₉₀	4.5	4.4	0.103	0.109	Mga	1.09	-0.7	9.62-	MgO	~				3
	A1203	0.33	7. 7.57	0.426	925.0	A1203	4.26	7.3	55.4	A1203	~			-	3
	Si02	17.8	17.6	0.293	0.324	ر د	3.24	2'7:	-177.6	2015	6				
	lotal	Ē.	100.0					-8.3	-349.3	Total	\$				
SuA2	CaO	27.7	21.7	0.387	0.387	Casio)	3.87	-6.2	.560.8	CaO	53	0.48	119	2.08	
	ÇĞ.	15.7	15.7	0.389	0.389	MgS 103	3.89	-5.6	-226.4	MgO	22				
	A1203	8.0	9.0	0.008	0.000	A1 203	9.08	0.0	9.	A1203	0				
	Si02	8.19	61.8	1.028	0.252	2018	2.52	1.4	59.0	2015	118				_
	lotal	0.031	100.0		,			-10.2	-427.2	Total	228				_
SAS(D)	Srō	56.5	57.3	0.553	0.413	Srsio3	4.13	1.01-	-421.7	SrO	107	0.40	26	1.96	_
	A1203	17.6	17.9	0.176	0.176	A1203	1.76	0.5	22.9	A1203	-				
	2015	24.4	24.8	0.413	0.140	Sro	1.40	-2.4	.tot.	Si02	۵				
	Total	98.5	100.0					-12.0	-500.2	Total	147				_
KHAS(A)	K20	16.5	16.5	0.175		K2S103	1.73	-7.3	-305.4	K20	•	0.36	7	1.15	
_	P B	13.0	2.0	0.323	_	HgS 103	3.23	-2.1	.188.0	OG M	•				_
	A1203	9. 9.	17.9	0.176	0.176	,A1203	1.76	5:0	52.9	A1203	_	_			
	2015	52.0	52.6	0.875	_	Sioz	9.73	2.1	204.8	5102	= :				
	Total	160.3	100.0					B.0-	7.692	lotat	3				-

TABLE 12

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-11	.vu	

	•	
Relevant oxides	(kcal/mole)	(kJ/mol)
$sio_2 + H_2o \longrightarrow H_2sio_3$	5.6	23.4
(vitreous silica)		
$Al_2O_3 + 3H_2O \longrightarrow 2Al(OH)_3$	3.1	13.0
$MgO + H_2O \longrightarrow Mg(OH)_2$	- 6.5	-27.2
$CaO + H_2O \longrightarrow Ca(OH)_2$	-13.1	-54.8
$sro + H_2o \longrightarrow sr(OH)_2$	-17.3	-72.4
$Na_2O + H_2O \longrightarrow 2NaOH$	-33.5	-140.2
$K_2O + H_2O \longrightarrow 2KOH$	-46.1	-192.9
$TiO_2 + H_2O \longrightarrow Ti(OH)_2O$	16.0	66.9
$P_{2}O_{5} + 3H_{2}O \longrightarrow 2H_{3}PO_{4}$	-55.9	-233.9
$B_{2}O_{3} + 3H_{2}O \longrightarrow 2H_{3}BO_{3}$	-9.8	-41.0
$\operatorname{ZrO}_2 + \operatorname{H}_2 \circ \longrightarrow \operatorname{Zr}(\operatorname{OH})_2 \circ$	-7.1	-29.7
Relevant Silicates		
$\text{Na}_2\text{SiO}_3 + 2\text{H}^+(\text{aq}) \longrightarrow \text{H}_2\text{SiO}_3 + 2\text{Na}^+(\text{aq})$	-28.8	-120.5
$K_2 sio_3 + 2H^+(aq) \longrightarrow H_2 sio_3 + 2K^+(aq)$	-41.7	-174.5
$MgSiO_3 + 2H^+(aq) \longrightarrow H_2SiO_3 + Mg^{2+}(aq)$	-13.9	-58.2
$Casio_3 + 2H^+(aq) \longrightarrow H_2Sio_3 + Ca^{2+}(aq)$	-16.1	-67.4
$Srsio_3 + 2H^+(aq) \longrightarrow H_2Sio_3 + Sr^{2+}(aq)$	-24.4	-102.1
$BaSiO_3 + 2H^+(aq) \longrightarrow H_2SiO_3 + Ba^{2+}(aq)$	-37.3	-156.1
$ZnSiO_3 + 2H^+(aq) \longrightarrow H_2SiO_3 + Zn^{2+}(aq)$	-2.4	-10.0
<u>Disilicates</u>		
$Ca_2SiO_4+3H^+(aq) \longrightarrow H_2SiO_3+2Ca^{2+}+OH^-$	-30.3	-126.8
$sr_2sio_4+3H^+(aq) \longrightarrow H_2sio_3+2sr^{2+}+OH^-$	-50.3	-210.5
$Mg_2SiO_4+3H^+(aq) \longrightarrow H_2SiO_3+2Mg^{2+}+OH^-$	-17.9	-74.9

Although calculations were undertaken assuming the presence of the simplest silicates choice of other silicates (such as disilicates) does not appear to change the calculations much. For example given below is the calculation for an assumed composition which shows only minor differences for the free energy of hydration calculated.

Composition (m	oles/	100g)	0.76	7 CaO		•	
	•		0 41	.7 MgO	•		
-			0.66	9 sio ₂		•	
0.384 Ca ₂ SiO ₄	=	-11.6	or	0.433	CaSiO ₃	=	-7.0
0.285 MgSiO ₃	· =	- 4.0		0.236	$MgSiO_3$	=	-3.3
0.132 MgO	=	- 0.9		0.334	CaO	=	-4.4
		-16.5		0.181	MgO	=	- <u>1.2</u>
							-15.9
•							
			or	0.384	Ca_2SiO_4	· =	-11.6
				0.132	${ m Mg}_2{ m SiO}_4$	=	- 2.4
				0.153	MgSiO ₃	=	- 2.1
						·	-16.1

The applicants have found that when the free energy of hydration is more negative than -10kcal/100 grams (-418.4kJ/kg) of composition the composition showed high solubility. The compositions where this relationship broke down were those for which the total solubility was not available (for example those materials containing sodium, where any dissolved sodium would be swamped by the sodium in the saline solution) or where the free energy of hydration of the most likely species present was not available from the literature.

As a test of this technique the two examples of European Patent No. 0399320 were examined. The disclosed examples had the compositions:-

Component	Composition 1 Weight percent	Composition 2 Weight percent
sio ₂	60.7	58.5
Al ₂ O ₃	-	5.8
CaO	16.5	3.0
MgO	3.2	-
B ₂ O ₃	3.3	11.0
Na ₂ O	15.4	9.8
K ₂ O	0.7	2.9
Iron oxide	0.2	0.1
BaO	-	5.0
ZnO	-	3.9

Using the above method of calculation Composition 1 had a free energy of hydration of -11.6kcal/100grams (-485.3kJ/kg) whereas Composition 2 had a free energy of hydration of -5.8kcal/100grams (-242.6kJ/kg). This would suggest that Composition 1 would be a saline soluble fibre, and hence physiologically safer than an insoluble fibre; whereas Composition 2 would be predicted to be a relatively insoluble fibre and hence less safe. This is what is disclosed in EP 0399320, the fibres of Composition 2 having a longer lifetime introduced the fibres were which in studies interperitoneally into rats.

As mentioned above this predictive test can fail under some circumstances. To avoid these difficulties the applicants looked to a different predictive technique, namely the assessment of the amount of non-bridging oxygens present. This is calculated by normalising the chemical analysis to 100 weight%; calculating the molar percentage of each oxide; summing the oxygen-weighted contribution of each oxide to get the total number of oxygens; summing the weighted contribution of each oxide of non-bridging oxygens (see below); and taking the ratio of non-bridging oxygens to the total number of oxygens. The applicants have found that when this figure exceeds 30% the fibres are soluble.

To explain the term non-bridging oxygen one must look to the structure of glasses. Glasses are very stiff liquids and to form usually require the presence of a material that can form a network (usually an oxygen-bridged network). The network may be modified by constituents that contribute non-bridging parts to the network and open the structure of the network and so prevent crystallisation. These materials are usually referred to as network-formers and modifiers respectively.

The terms modifier and network former are well known in the glass industries. Network formers are materials such as \sin_2 , P_2O_5 , B_2O_3 and GeO_2 which can form an interconnected network to form the glassy phase. Modifiers are substances such as CaO, Na_2O , and K_2O which alter the network and have effects on such properties as viscosity and melting point. There are some intermediate materials (such as Al_2O_3 , TiO_2 , PbO, ZnO and BeO) which can act as both network formers and modifiers depending on the environment and on the amount present.

In the above mentioned test, for calculating the non-bridging oxygens, one ignores the network formers and calculates the contribution of each other oxide. The contribution of each oxide depends on the geometry and charge of each cation in the glass. As examples typical contributions are as follows:-

 ${\tt Ca}^{2+}, {\tt Mg}^{2+}, {\tt Sr}^{2+}$ and other divalent network modifier cations contribute 2 non-bridging oxygens ${\tt K}^+, {\tt Na}^+$ and other monovalent network modifier cations contribute 1 non-bridging oxygen ${\tt Al}^{3+}, {\tt Ti}^{3+}$ and other intermediate cations contribute -1 non-

Al³⁺,Ti³⁺ and other intermediate cations contribute -1 nonbridging oxygen (i.e these oxides <u>reduce</u> the number of nonbridging oxygens)

 $({\rm Ti}^{4+} \ {\rm is} \ {\rm reduced} \ {\rm to} \ {\rm Ti}^{3+} \ {\rm in} \ {\rm most} \ {\rm glasses} \ {\rm when} \ {\rm present} \ {\rm in} \ {\rm relatively} \ {\rm small} \ {\rm quantities})$

38

TABLE 13

																					_		_			_	_	_		_	<u> </u>	_	_		-	_		_	Т		_	
Log(norm	solub)	1.63				1.82				1.88				1.84				2.40			7.7		9	0.0				0.0 				9.18	-		8	3	_		2 16	۲.45		
Hormal ized	Solubility	43				3				92				\$				549		277	<u>co</u>			*				œ —				ه —			-	, 			200	700		
	m2/9	0.30				0.39				0.36				0.40				0.37		3, 6	6,40		,	9.70				0.62				0.52			9			_		3 .		
įίχ	(mdd)	77	0	~	15	28	_	57	103	ž.	0	8	110	29	•	67		~	367	300	- (S :	294	2	•	· ·	6	= ·	4	ın (02	~ •	,	• 5	2	٠.	۷ ۱	٠.	ا ا	٥;	301	460
Solubility	Oxide	CaO	A1203	2018	Total	080	A1203	2015	Total	C80	A1203	2018	Total	CaO	1 102	2015	Total	oro.	2015	1000	ָרָפּק ביים	sio2	Total	No.20	A1 203	\$102	Total	CaO	A1203	2015	Total	CBH	A1 203	2016	1000	o lo	A1203	2015	lotal	Sro	200	fotel
2M.B.O		48.8				48.4				50.4				55.8				8.89		-	7.40	ę		-5.5				6.7				-5.4				<u>:</u>				65.8		
N.B.O.		78.4				77.8	_			9.08				85.1				102.4			2.79			-12.6				3.8				-11.2				-2.0				o.&		
Oxygen	Total	160.8				160.9				159.6				152.5				148.8.			151.4			228.3				198.1				202.6				3.102				150,5		
Hol.x		54.9	15.7	29.4	100.0	24.0	15.1	30.8	100.0	55.4	15.2	29.3	100.0	47.5	6.6	42.6	100.0	51.2	48.8	100.0	48.6	51.4	100.0	22.0	28.3	49.7	100.0	25.7	23.8	50.5	100.0	17.0	22.6	5.09	100.0	23.9	25.3	50.8	100.0	23.5	26.0	100.0
Holes		0.842	0.240	0.451	1.533	0.829	0.232	0.473	1.534	0.852	0.234	0.451	1.537	0.761	0.159	0.682	1.602	0.609	0.581	1.190	0.822	0.870	1.692	0.303	0.389	0.684	1.376	0.374	0.347	0.736	1.457	0.248	0.331	0.882	1.461	0.288	0.305	0.612	1.205	0.338	0.374	0.726
tion	V.X	47.2	24.5	1.72	Total	5.97	23.7	28.4	Total	47.8	23.9	27.1	Total	42.7	12.7	41.0	Total	63.1	34.9	Total	46.1	52.3	Total	18.8	39.7	41.1	Total	21.0	35.4	44.2	Total	10.0	33.7	53.0	Total	29.8	31.1	36.8	Total	35.0	21.0	43.6 Total
Composition	Ox ide	Ca	A1203	5102		Sa	A1203	5102		3	A1203	5102		Crio	1102	5102		SrO	5102		Ca0	5102		Na20	A1203	\$102		CaO	A1203	2018		PB0	A1203	Si02		Sro	A1203	5102		Sro	CaO	s 102
Helt Code	÷	CASIOCAL - A	:			CAS10(B) - B				CAS10(B) - C				C15(A)				SrSi03			CaS 103			NAS(A)				CAS4				MAS(B)				SAS(A)				SCS(A)		

TABLE13 (continued)

_									_			_	_	_		_	<u> </u>						_		_					1			_	1				٦
Log(norm	solub)	!					1.98					2.33					1.72				1.38					2.08				70	:			1.15				7
Normal ized	Solubility	38					95					212					52		•	•	77					119				8	!			21				
Γ	m2/g	١					0.39					0.39					0.38				0.30					0.48				07 0	}		0	0.36				
l ty	(bbm)	25	s	~	'n	69	110	55	_	22	148	147	19	0	123	331	02	_	8	79	=	۷	m	80	29	53	25	0	81.6	107	-	3	147	78.	6 0	_	= 5	۷,7
Solubility	Oxide	SrO	Mg0	At 203	Sioz	Total	Sro	MgO	A1203	5018	Total	Sro	N ₉ 0	A1203	2018	Total	Sro	. A1203	Sio2	Total	CoO	MgG	A1 203	sioz	Total	CaO	₩30	A1203	2015	Sro	A1203	\$102	Total	K20	MgO	A1 203	Si02	10101
XH.B.0		35.8					53.8					50.5					36.3				23.1					24.0				20.7	;			23.1		٠.		-
₩.8.0.		8.09					84.8					9.08					61.4				41.4					85.0				6,43	!			5.15				
Oxygen	lotal	169.6					157.6					159.5					169.3				179.3					157.5				0.491				179.5				
Hol.X		49.1	8.8	21.5	56.6	100.0	9.33	10.0	12.2	33.2	100.0	20.7	20.2	0.6	58.4	100.0	47.7	17.0	35.3	100.0	42.6	9.2	29.5	20.3	100.0	21.4	21.5	0.4	790.	2.8.5	15.4	36.1	100.0	11.3	20.8	11.4	56.6	100.0
Holes		0.531	0.030	0.232	0.288	1.081	0.531	0.119	0.145	0.394	1.189	0.318	0.310	0.010	0.895	1.533	0.540	0.193	0.399	1.132	0.624	0.112	0.432	0.296	1.464	0.387	0.389	0.008	1.028	0.545	0.173	0.40	1.124	0.175	0.323	0.177	0.879	1,239
ition	Mr.X	55.0	1.2	23.7	17.3	Total	55.0	4.8	14.8	23.7	Total	33.0	12.5	1 .0	53.8	Total	56.0	19.7	24.0	Total	35.0	4.5	0.22	17.8	Total	21.7	15.7	8.0	01.0	5.45	17.6	54.4	Total	16.5	13.0	18.0	52.8	
Composition	0x ide	Sro	Hg0	A1203	2015		Sro	Mg0	A1203	Si02		Sro	MgO	A1203	Si02		SrO	A1203	Si02		Ca0	Mg0	A1203	Sio2		CaO	S .	A1203	2015	62.0	A1203	5102		K20	MgO	A1203	S102	
Helt Code		SHAS(A)					SHAS(B)					SHS(A)					SAS(C)				CHAS(A)					SNAZ				SAS(D)				KHAS(A)				

Table 13 shows, for the compositions of Table 11, the calculated figure for non-bonding oxygens and these figures are reproduced in Fig.5. It can be seen that the plot of Fig.5 is more linear than that of Fig.4. Compositions with greater than 30% non-bridging oxygens show high solubility.

To illustrate the method the calculation for one example (the first shown in Table 13) is given below:-

	Mols/	Total Oxygen	Non-bridging Oxygens
CaO Al ₂ O ₃ SiO ₂	0.549 0.157 0.294 	0.549 0.471 (0.157*3) 0.588 1.608	1.098 (2*0.549) -0.314 (-1*0.157*2Al) 0.000

% non-bridging oxygens (0.784/1.608)*100 = 48.8%

As an example one can look to European Patent Specification No. 0399320 referred to above. Using this method in relation to that specification Composition 1 has a non-bridging oxygen percentage of 48.2% whereas Composition 2 has a non-bridging oxygen percentage of 19.6%, again predicting that Composition 1 is more soluble than Composition 2.

There is a further criterion which a composition must meet for it to form vitreous fibres, namely that it be capable of forming a glass. The applicants have found a simple test. If the ratio of modifiers to network formers is less than a critical value (for SiO₂ based glasses, 1) the composition will generally form a glass. For the purpose of this test reasonable results are obtained if such intermediate materials are treated as networkers. Table 14 shows for a series of compositions in each column:-

Fibre reference (N.B. these are not the same fibres as shown in tables 9 & 10)

Composition

Molar ratios

Ratio of glass modifiers to network formers

Free energy of hydration of raw oxides

Melting Point

X-ray diffraction results

Solubility data (# indicates not measured)

Specific surface area

Normalised solubility data

Arbitrary ranking as to solubility and glass forming ability

Indication as to whether melting point above 1500°C

It should be emphasised that this test is a screening one rather than a wholly predictive one as there are several circumstances that may lead to its failure. Among these circumstances are compound formation and inability to quench fast enough to form glass.

Having adopted these tests as a screening method there follows a further step to ascertain whether the composition will form a vitreous fibre. This last step is best examined experimentally as fibre forming ability is a complex function of many physical characteristics, e.g. viscosity, which are often difficult to measure.

Key to Fig.1

CR Cristobalite

TR Tridymite

PS Pseudowollastonite

WO Wollastonite

RA Rankinite

LI Lime

PE Periclase

FO Forsterite

PR Protoenstatite

DI Diopside

AK Akermanite

ME Merwinite

MO Monticellite

TABLE 14

							1 F	BL	ت.	1	4												
1200C	_		-	F	-	Ī										-		-		it .	-	E	
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58100	Ē			<u> </u>								Ī	-	=								=	
7 107 8 10 S	Ê		E .	Ē	47.	-	503	17.			,	Ī	42	43		7		=		7		23	
	Ĕ		Ē		ŕ	- 4	3,55	3				ĺ	201	124		ا			Ē -	-	5	<i>(</i> ()	
Z.X.	-		ZS:0	_U:57	0:53	0:53	ם, פון	12.0	5	7	;	Ê	10:01	ř	1	5 9	;	/A:D	U:72	19:0.	ł	/S:n	
10131	I	è	162	** **	8	Ê	\$2		Š	-	227	٦	375		,	-	ž	20	~	24 107		200	
S Unys	113	125	872	7	200	Ť,	Ž	7		~~	· 0	ĺ	333	ļ	7		22					285	
Data (par		,2, 12,	9	2	25		•			, ,	Š	ř) S	1	3	<u> </u>	18 V	200	2	2 2	×===	222	
	 	~%	2°2			<u> </u>	[2	1265	2	298	325	<u> </u>	595	<u> </u>	<u> </u>			1	1	<u> </u>		1	<u> </u>
Soluti	٥	1201 5102	200	A 205	202	500	2012	200	205	V 203	50.5		7320 C00	2025	200	2 2 2 2 2 2 2	A1205	41205 61205	200	202	S10 S10 S02 S02 S02 S02 S02 S02 S02 S02 S02 S0	2005	
MRD 07165	Amin [7] 50: T 11 21 8	leaks)	(No proks)	Alprijstour Trace	iclossy troic)		Amor Aribus (strang) some crystel. (strang)	(3 small peaks)	some glass (strong)	1 1 20	(No glass)	Trytatine Table	Crystalling a	Amortaions cryst.)	Some crystal.		28	Andi fricis	Angi (Nous (No praks)	Aliasi jaiotis (No peaks)	(No peaks)	(2 small peaks)	
 		C. M.	(fut.)	1400 C)		1365 (Fut.)	1543		(filt.)	5_	(Cut.)	231	fert. 1	157	(5.16.)	1565 (Eut.)	(Eut.)	1541	(3,0051)	(1400 (2)	1670	E (101)	
(kcel/mot le		4	=	3.0	.6.1	1:32:	1.4.	,		2.1.	.4.3	5:A	.8.			,2.	7.	.2.1		3.6	6	3.6	
Hediller/	0.D		B. C	Z.0.	49:0	1:35				0		:570;	, 1.1gr	17.	1		0.41	0.31	72:0	12:0	12:0	A:0	-
: 5	P	1.0 A1201 1.9 S102	~0.	10.0 10.0 10.0 10.0 10.0 10.0 10.0 10.0	200	:20	==	<u> </u>	==	===	==			~::	~~	<u>~=</u>	P.~	2		N.P.B.	Pos	000	1
Gigs StigitHets		~~	~04	2.4 A 205	29: \$ \$1203 39: 2 \$1203		00	2	200		200		10 K R320	86	-	_~	D-4		34.0 A1205	-		17.5 Cad	7
10 2 MB 12:21	÷	~~	10. 19	HAS(A)	MAS(A) 21	Ī		CBS105 62		18.28 / S	l							MASTER	EAST. NO. 34	HASIES	SAS(A)	SCS(A)	

43 TABLE 14 (continued) 22 260 (1500 C) (3 soult peaks) Anor prous (No proks) (No peaks) Anoi mous (No pents) (C) 15th (Eut.) 1580 C. (Eut.) (Evit) 1501 151111 Eu. . E Euc ? 8 -€ 17.7 Ē 9 7.0 50. [AAS[A] --74. C7 542.67 50. C8 HAS(E) SKS103 -

CLAIMS

1. A method of predicting whether a given composition will form a saline soluble vitreous material, the method comprising the steps of:-

EITHER

- a) assuming the composition to be a mixture of the compounds that would be present at equilibrium (on the basis of knowledge, informed belief or reasonable assumption);
- b) calculating the sum of the free energies of hydration of each of the compounds assumed to be present;
- c) if the sum of the free energies of hydration is more negative than -10 kcal/100grams (-418.4kJ/kg) of composition, assuming the composition to be soluble in saline solution;

OR

- a") calculating the percentage of non-bridging oxygens present in relation to the total oxygen present and, if the percentage is greater than 30% assuming the composition to be soluble in saline solution;
- 2. A method as claimed in claim 1 which further comprises in combination with steps a) -c) and/or a"):-
- d) calculating the ratio of glass modifier to network former present;
- e) if the ratio is less than a critical value (for SiO₂ based compositions, 1) assuming that the composition will form a vitreous material.

- 3. A saline soluble vitreous inorganic fibre selected and used in the knowledge that it has a composition meeting the criteria that:-
- a) the calculated sum of the free energies of hydration of the compounds that would or could be present at equilibrium (on the basis of knowledge, informed belief or reasonable assumption) is more negative than -10 kcal/100 grams (-418.4kJ/kg) of composition;

and/or

- b) the percentage of non-bridging oxygens present in relation to the total oxygen present is greater than 30%.
- 4. A saline soluble vitreous inorganic fibre as claimed in claim 3 selected and used in the knowledge that it has a composition meeting the criterion that the ratio of glass modifiers to network formers is less than a critical value (for SiO_2 based compositions, 1).
- 5. Use of a saline soluble vitreous inorganic fibre in the knowledge that it has a composition meeting the criteria that:-
- the calculated sum of the free energies of hydration of the compounds that would or could be present at equilibrium on the basis of knowledge, informed belief or reasonable assumption is more negative than -10 kcal/100grams (-418.4kJ/kg) of composition;

and/or

- b) the calculated percentage of non-bridging oxygens in relation to the total oxygen content is more than 30%.
- 6. Use as claimed in claim 5 of a saline soluble vitreous inorganic fibre in the knowledge that it has a composition meeting the criterion that the ratio of glass modifiers to network formers is less than a critical value (for SiO₂ based compositions, 1).

7. Use as saline soluble fibres having a shrinkage of less than 3.5% when exposed to 1000°C for 24 hours and/or 800°C for 24 hours, of vitreous fibres having a composition comprising (in weight %):-

```
SiO_2 >58% - (for MgO =< 10%) and SiO_2 >58% + 0.5(%MgO -10) - (for MgO >= 10%) CaO 0% - 42% MgO 0% - 31.33% Al<sub>2</sub>O<sub>3</sub> 0% - <3.97%
```

and being essentially free of fluxing components such as alkali metals and boron oxide.

- 8. Use as claimed in claim 7 of saline soluble fibres having a shrinkage of less than 3.5% when exposed to 1000°C and/or 800°C for 24 hours, of vitreous fibres in which the amount of SiO₂ is <70%.
- 9. Use as claimed in any of claims 7-8 of saline soluble fibres having a shrinkage of less than 3.5% when exposed to 1000°C and/or 800°C for 24 hours, the fibres crystallising as diopside and having the composition consisting essentially of:-

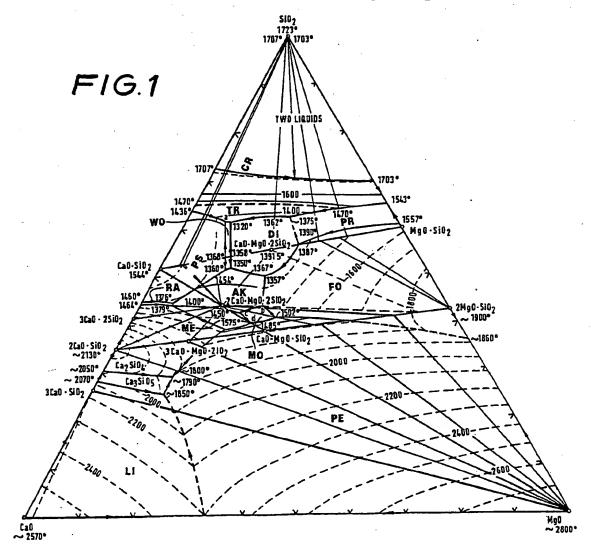
Component		Composition A
		Weight percent
sio ₂	•	59-64
Al ₂ O ₃	•	0-3.5
CaO		19-23
MgO		14-17
		14-17

10. Use as claimed in any of claims 7-8 of saline soluble fibres having a shrinkag of less than 3.5% when exposed to 1000°C and/or 800°C for 24 hours, the fibres crystallising as wollastonite and/or pseudowollastonite and having the composition consisting essentially of:-

Component	Composition B
	Weight percent
sio ₂	60-67
Al ₂ 0 ₃	0-3.5
CaO	26-35
МдО	4-6

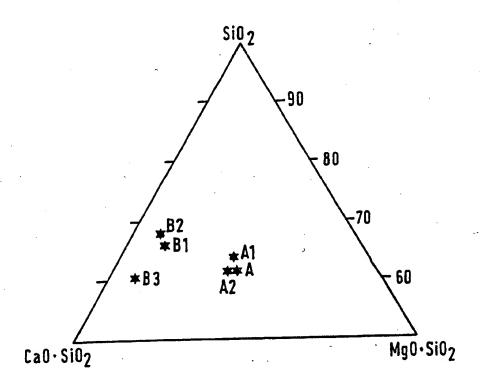
- 11. Use as a saline soluble fibre as claimed in any of claims 5-10 and in which, owing to the increase in molar volume on crystallisation, the linear shrinkage at the maximum service temperature is less than 3.5%.
- 12. A method for the manufacture of refractory oxide fibres containing calcium and silicon by the formation of an oxide melt containing calcium and silicon characterised in that all or part of the calcium and all or part of silicon is provided by a calcium silicate.
- 13. A method as claimed in claim 13 in which the calcium silicate is wollastonite.

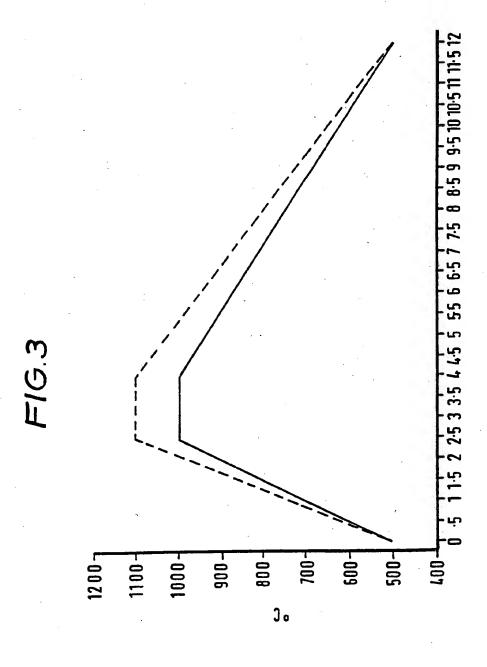
CR) TR}·····	SiO ₂
PS	∝-CaO·SiO2
	ß-(CaMg)0·SiO ₂
RA	3 Ca0 · 2 Si0 2
LI	(CaMg)O
P.E	Mg0
	2(Mg,Ca)0.Si02
PR	(Mg, Ca)0·SiO2
DI	(Ca, Mg)0.Mg0.2Si02
AK	2Ca0·Mg0·2Si02
ME	3Ca0 · Mg0 · 2Si02
M0	(Ca,Mg)O·MgO·SiO2

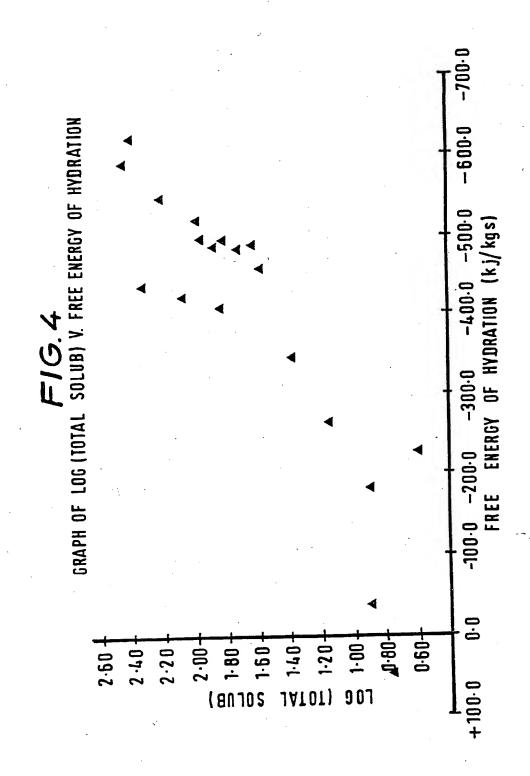


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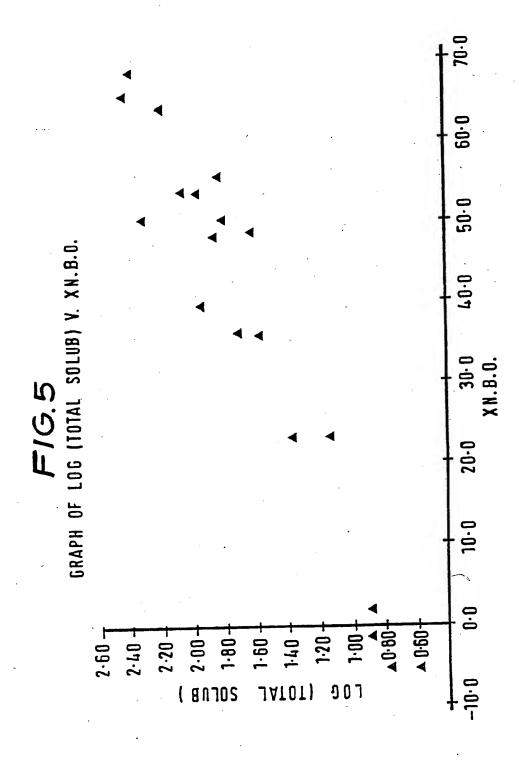
FIG.2











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		Documentation Searched other to the Extent that such Documents	than Minimum Documentation are Included in the Fields Searched ⁸	
III. DOCUM		D TO BE RELEVANT 9		1 5
Category ⁶	Citation of D	ocument, 11 with indication, where appropri	ate, of the relevant passages 12	Relevant to Claim No. ¹³
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Date of the	-	the International Search	Date of Mailing of this International Se	earch Report
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The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

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